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24 November 2026





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Open and Closed Radicals: Local Geometry Around Unpaired Electrons Governs MAS DNP Performance

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KEYWORDS: MAS DNP, cross-effect, nitroxide radicals, radical design, open, closed

ABSTRACT: The development of magic-angle spinning dynamic nuclear polarization (MAS DNP) has allowed atomic-level characterization of materials for which conventional solid-state NMR is impractical due to the lack of sensitivity. The rapid progress of MAS DNP has been largely enabled through the understanding of rational design concepts for more efficient polarizing agents (PAs). Here, we identify a new design principle which has so far been overlooked. We find that the local geometry around the unpaired electron can change the DNP enhancement by an order of magnitude for two otherwise identical conformers. We present a set of 13 new stable mono- and di-nitroxide PAs for MAS DNP NMR where this principle is demonstrated. The radicals are divided into two groups of isomers, named open (O-) and closed (C-), based on the ring conformations in the vicinity of the N-O bond. In all cases, the open conformers exhibit dramatically improved DNP performance as compared to the closed counterparts. In particular, a new urea-based biradical named HydrOPol and a mono-nitroxide O-MbPyTol yield enhancements of 330 ± 60 and 119 ± 25 at 9.4 T and 100 K respectively, which are the highest enhancements reported so far in the aqueous solvents used here. We find that while the conformational changes do not significantly affect electron spin-spin distances, they do affect the distribution of the exchange couplings in these biradicals. Electron Spin Echo Envelope Modulation (ESEEM) experiments suggest that the improved performance of the open conformers is correlated with higher solvent accessibility.

Dynamic nuclear polarization (DNP) is a consensus strategy to overcome the well-recognized sensitivity limitations of nuclear magnetic resonance (NMR) experiments.[1-3] In particular, DNP induced sensitivity enhancements have been demonstrated in magic-angle-spinning (MAS) NMR experiments on materials samples[4-5], such as polymers,[6-10] porous and structural materials[5, 11-16], nanoparticles[17-19], pharmaceuticals[20-23], zeolites[24-25], catalysts[26-33], and for biomolecular samples[34-41] in 1D and 2D solid-state NMR experiments that are otherwise difficult or unfeasible, even with isotope labelling. DNP in liquid-state experiments has also been shown to yield very substantial ^{13}C enhancements [42-46].

In DNP experiments, the comparatively high electron polarization is transferred to nuclear spins via microwave irradiation at or close to the EPR transition. The applied microwave field saturates electron spin transitions that are hyperfine-coupled to nearby nuclear spins. Thereby, the more strongly polarized electron spins can exchange populations with nuclear spins. Several possible mechanisms mediating the transfer have been identified: *solid effect (SE), cross effect (CE), thermal mixing (TM) and Overhauser effect (OE)*.[47-51]The mechanism at play depends on the magnetic field strength, temperature, nuclear and electron spin concentration, and the type of polarizing agent (PA). The source of unpaired electrons is vital for efficient DNP, and has been the subject of intense recent scrutiny[52]. With the CE mechanism being typically the most efficient at high magnetic fields and temperatures around 100 K, bi-radicals have so far proved to be the most efficient polarization source for MAS DNP.[53-60] The best performing PAs for high-field DNP are either di-nitroxides[54, 58, 61], or mixed nitroxide-narrow line radical systems[53, 60, 62-63]. High-spin transition metal complexes have also been demonstrated as MAS DNP PAs.[64-65]

Despite this intense development, the best performing radicals currently deliver ^1H DNP enhancements of around $\varepsilon = 300$ at 9.4 T and 100 K in bulk frozen solutions[54-55, 66], which is significantly less than the theoretical maximum (corrected for depolarization[67-68]) of

~658 for ^{1}H ($|\gamma_e/\gamma_n|$). By lowering the temperature to 55 K or by improving the microwave penetration by adding dielectric particles, enhancement factors close to the theoretical maximum can be obtained with bisnitroxides.[69,70]

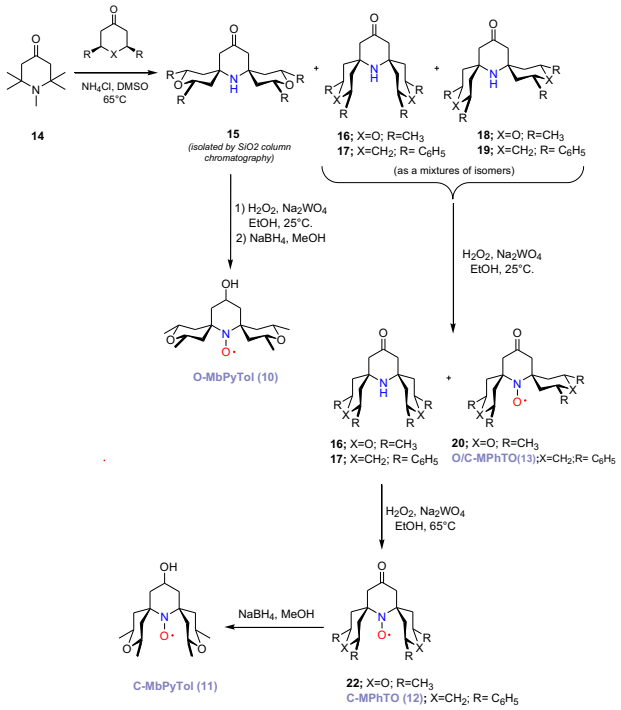
Over the last decade, a number of design criteria for efficient PAs have been identified and implemented. Initially, TEMPO was used due to its availability, providing a proton enhancement of ~40 at 9.4 T at 100 K. In 2004, the use of bi-radicals with limited flexibility and *well-defined inter-electron distances*, leading to large dipolar couplings, led to the introduction of the BTnE series (^{1}H ε~175 at 84 K)[71] and TOTAPOL (^{1}H ε~70 at~100 K)[56]. In 2009, bTbK was introduced[72], (^{1}H ε~250 at 94 K and 5 T) which has two TEMPO moieties linked by a rigid tether, leading to a *well-defined relative orientation of the nitroxide moieties.* This geometrical arrangement translated to more efficient CE matching[2-3] and its theoretical understanding has been recently developed.[73] A *large electron-electron dipolar coupling*, essential for CE, is ensured by an average interelectron distance of ~11.8 Å in this class of radicals[72]. In 2012, *long electron relaxation times* were identified as a design criterion, yielding the introduction of bulky substituents on the nitroxides in the bCTbK[21, 74] (^{1}H ε~105 at ~100 K and 9.4 T in 1,1,2,2-tetrachloroethane (TCE) and ^{1}H ε~60 at 105 K and 9.4 T for microcrystalline glucose and sulfathiazole) and TEKPol (^{1}H ε>200 for16 mM in TCE at 105 K and 9.4 T) families[74-75 55], which are suitable for organic solvents, followed by the water soluble PyPol (^{1}H ε~200 for 10 mM in water/glycerol mixture at 97 K and 9.4 T) and AMUPol (^{1}H ε~230 for 10 mM in water/glycerol mixture at 97 K and 9.4 T and ^{1}H ε~400 at 80 K and 8.9 T[76]) families in 2013[66]. In addition, hybrid systems consisting of an isotropic narrow EPR line radical covalently tethered to a broad-line nitroxide moiety have been demonstrated to preserve high enhancements at higher magnetic fields.[53, 63]

Here, we introduce and demonstrate the importance of a new design parameter: the *local geometry around the unpaired electron*. We show that DNP performance is dramatically affected, both at 9.4 T and 21.15 T, by changes in local conformation around the unpaired electron in mono- and di-nitroxides, for otherwise identical constitution irrespective of the radical concentration. We control the conformation of a tetrahydropyran ring substituent by including locking methyl and phenyl groups. We show that the bi-radical HydrOPol and the mono-radical O-MbPyTol yield enhancements of 330 ± 60 and 119 ± 25 at 9.4 T and 100 K in DMSO/water mixture. We use pulsed EPR to identify that, while keeping the mean spin-spin distances and relaxation properties nearly the same, changes in the local geometry lead to differences in solvent accessibility to the electron spin (which is also supported by measured reduction kinetics of the different mono-radicals), as well as to differences in the isotropic exchange coupling distributions in bi-radicals, and in turn correlate with the large differences in DNP performances experimentally observed.

EXPERIMENTAL

NMR spectroscopy. All DNP experiments were performed on a commercial Bruker Avance III 400 MHz NMR spectrometer equipped with a 263 GHz gyrotron microwave source using a 3.2 mm triple resonance CP-MAS probe at sample temperatures around 100 K with MAS at 8 kHz. All of the radicals were synthesized as reported below and in the SI. A 3.2 mm sapphire rotor, to optimize microwave penetration, was filled with 22 μL of radical containing solution. The solution was confined with a silicone plug to prevent any leakage and the rotor closed using either a zirconia or vespel cap. The solvents used were glycerol-d_8:D_2O:H_2O (60:30:10 v/v/v);[77] DMSO-d_6:D_2O:H_2O (60:30:10 v/v/v);[78] or TCE,[21, 79] commonly used in MAS DNP. Three freeze-thaw cycles were always used to reduce the presence of dissolved oxygen when using TCE,[55] with each cycle consisting in the insertion and ejection of the rotor into the stator. The microwave power was optimized for each sample to between 4 and 12 W, at the probe entrance, to obtain the largest DNP enhancements. The ^{1}H DNP enhancements of the solvent were measured through ^{1}H–^{13}C cross-polarization (CP) ($\varepsilon_{C\,CP}$) with a standard ramped (90%-100% or 70%-100%) cross polarization pulse sequence.[80-81] In some experiments in which the silicone plug was avoided, the enhancement has been directly measured with a standard proton echo acquisition after a presaturation loop. Details are given in Table S1. In all cases the integrated intensities of the solvent peaks, with and without microwave irradiation, were compared to determine $\varepsilon_{C\,CP}$ in order to account for any line narrowing arising from microwave induced sample heating.

EPR spectroscopy. The deuterium ESEEM measurements for assessing water accessibility[82-83] were performed on a Bruker Elexsys E580 EPR spectrometer equipped with a Bruker MS3 split-ring resonator at frequencies of approximately 9.3 GHz. The temperature of 50 K was achieved by liquid helium cooling. A 3-mm O.D. capillary loaded with roughly 60 μL sample was inserted into the resonator after shock-freezing in liquid nitrogen. Three-pulse ESEEM measurements consist of the following sequence of pulses: $\pi/2$-τ- $\pi/2$-T-$\pi/2$-τ-echo. The longitudinal electron magnetization is converted into transverse magnetization on allowed electron spin and nominally forbidden electron-nuclear spin transitions by the first $\pi/2$ pulse. During the first delay τ, electron transverse magnetization dephases and the second $\pi/2$ pulse flips most of the magnetization back to the z axis. Part of the magnetization is converted to transverse magnetization on nuclear transitions. Longitudinal relaxation and transverse nuclear spin relaxation occur during time T and the final $\pi/2$ pulse converts longitudinal electron spin magnetization and transverse nuclear magnetization into transverse magnetization on allowed and forbidden transitions. An echo signal is observed at time $T+2\tau$. The value of τ=344 ns was used in order to suppress ^{1}H modulations on the ESEEM decay envelope. The second inter-pulse delay, T, with an initial value of 80 ns was incremented in steps of 8 ns. The integrated echo intensity was measured as a function of T increment, with an integration gate of 32 ns length. The pulse lengths were 16 ns for the $\pi/2$ pulse and 32 ns for the π pulse.



Scheme 1: Synthesis of the O-, C- and O/C-isomers of MbPyTol and MPhTO nitroxides**.**

Echo-detected (ED) EPR measurements, as well as longitudinal and transverse relaxation measurements were performed at W band (94.1 GHz) on a Bruker Elexsys E680 spectrometer at 105 K. Transverse relaxation data were acquired with a Hahn echo sequence $\pi/2$-*t*- π-*t*-echo and applying a (+x, -x) phase cycle to the π /2 pulse. Longitudinal relaxation data were acquired with an inversion recovery pulse sequence: π-*T*-$\pi/2$-*t*-π-*t*-echo with a (+x, -x) phase cycle on the $\pi/2$ pulse. The pulse lengths were 12 ns for the $\pi/2$ pulses, 24 ns for the π pulse of the detection echo subsequence and for the inversion pulse. The initial values of 1 µs for *T* and 400 ns for *t* were used.

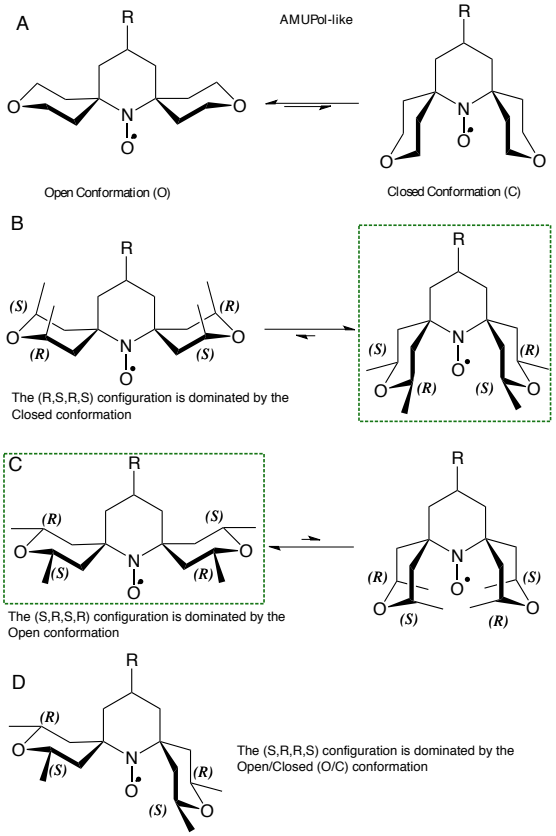
Molecular dynamics simulations. In order to correlate molecular structures with the DNP efficiencies, molecular dynamics simulations were carried out in explicit water for the O-radicals (**2**) and (**7**) and their C- counterparts (**3**) and (**8**) as described in detail in SI 9. The simulations were done for 300 ps trajectories in a cubic box with about 1400 water molecules using the GROMACS package using an AMBER force field (ff99SB).

Synthesis of the radicals. Experimental details of synthetic procedures and characterization of all compounds are given in the SI. The mono-nitroxides were prepared according to Scheme 1 starting from the corresponding pure cis 2,4,10,12-tetramethyl-7-aza-3,11-dioxadispiro[5.1.5.3]hexadecane-15-one isomers 15, 16 and 18, previously obtained by reacting 1,2,2,6,6-pentamethylpiperidine (14) with cis-2,6-dimethyltetrahydropyran-4-one in the presence of ammonium chloride via crossed aldol condensation (Scheme 1). During the reaction the three isomers, namely open (O-), closed (C-) and open/closed (O/C), are formed and isolated by column chromatography either at the amine stage or after oxidation to the corresponding nitroxide (see SI). The same procedure was followed the preparation of 17 and 19 but using cis-3,5-diphenylcyclohexanone. The conformations have been determined by X-Ray structure analysis of compounds 11, 12, 13, 15, 16, 17, 21, 22 (Figure S16-S19). The bisnitroxides were prepared using the previously reported procedures[54, 66] starting from mono-nitroxides 20-22.

RESULTS AND DISCUSSION

Open and closed conformers

Figure 1 shows all the radicals investigated in this work. In AMUPol[66] (**1**), the spirotetrahydropyran rings at the α-position of the nitroxide moiety can exist in two different chair conformations: pointing away from the NO• bond (denoted as open O-) or pointing toward the NO• bond (denoted as closed C-). Given the energy barrier for interconversion and the relative energies of the two states, AMUPol exists in solution as a mixture of these two populations of conformers (Scheme 2).



Scheme 2: Conformational analysis: A) conformational equilibrium active in each nitroxide moiety of the AMUPol biradical. By introducing *cis* 2,6-dimethyl groups in the tetrahydropyran ring, is possible to lock the conformational equilibrium, favoring the conformation where the ring substituents assume equatorial positions. Reading clockwise the absolute configuration in panels B), C), D) we show that the (*R,S,R,S*) configuration corresponds to a Closed conformation (panel B), the (*S,R,S,R*) configuration to an Open form (panel C) and the (*S,R,R,S*) configuration to a mixed Open/Closed conformation (panel D).

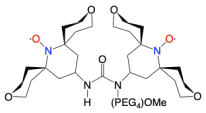
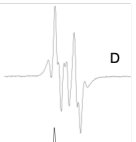
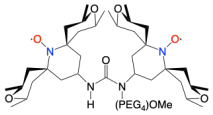
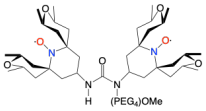
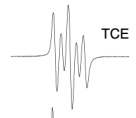
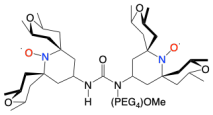
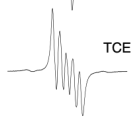
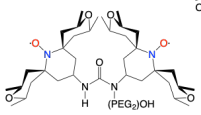
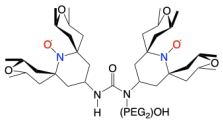
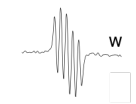
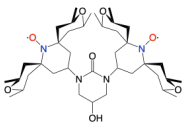
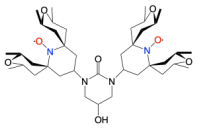
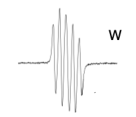
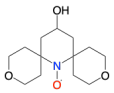
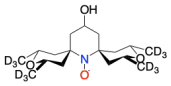
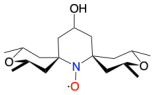
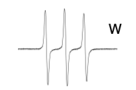
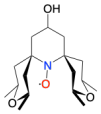
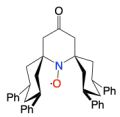
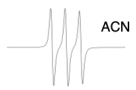
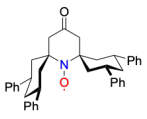
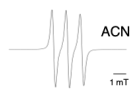
Radical	Structure	EPR spectrum	EPR parameters	$\varepsilon_{on/off}$ DMSO-d_6:D_2O:H_2O (60:30:10 v/v/v)	$\varepsilon_{on/off}$ TCE	$\varepsilon_{on/off}$ Gly-d_8:D_2O:H_2O (60:30:10 v/v/v)
AMUPol (726 g·mol^{-1}) (**1**)		 D	|J| = 45 ± 2 MHz* A_{iso}= 46.7 ± 0.2 MHz	175 ± 20 (8.2 ± 0.8 s) (3.2 mM)		244 ± 20* (3.3 ± 0.5 s) (10 mM)
O-MAMUPol (839 g·mol^{-1}) (**2**)		 TCE	|J| = 48 ± 2 MHz $\sigma(J)$= 8 ± 3 MHz A_{iso}= 44.7 ± 0.2 MHz	250 ± 20 (7.0 ± 0.7 s) (3.2 mM)	102 ± 20 (1.5 ± 0.7 s) (16 mM)	
C-MAMUPol (839 g·mol^{-1}) (**3**)		 TCE	|J| = 0 ± 50 MHz $\sigma(J)$= 150 ± 50 MHz A_{iso}= 43.0 ± 0.2 MHz	23 ± 3 (5.5 ± 0.7 s) (3.2 mM)	14 ± 2 (0.9 ± 0.1 s) (16 mM)	
O/C-MAMUPol (839 g·mol^{-1}) (**4**)		 TCE	|J| = 110 ± 5 MHz $\sigma(J)$= 12 ± 5 MHz A_{iso}= 44.7 ± 0.2 MHz	211± 16 (21.0 ± 2.0 s) (2.0 mM)	78 ± 3 (1.1 ± 0.7 s) (16 mM)	
HydrOPol (736 g·mol^{-1}) (**5**)		 W	|J| = 42.0 ± 2 MHz $\sigma(J)$= 0.5 ± 0.2 MHz A_{iso}= 44.3 ± 0.2 MHz	330 ± 60 (3.8 ± 0.5 s) (10 mM)		
C-HydrOPol (736 g·mol^{-1}) (**6**)		 W	|J| = 450 ± 50 MHz $\sigma(J)$≦100 MHz A_{iso}= 44.3 ± 0.2 MHz	29 ± 3 (1.9 ± 0.7 s) (10 mM)		
O-PyPOLC$_6$OH (704 g·mol^{-1}) (**7**)		 W	|J| = 46.6 ± 2 MHz $\sigma(J)$= 4 ± 3 MHz A_{iso}= 46.7 ± 0.2 MHz	203 ± 15 (3.6 ± 0.4 s) (10 mM)		
C-PyPOLC$_6$OH (704 g·mol^{-1}) (**8**)		 W	|J| = 450 ± 50 MHz $\sigma(J)$≦100 MHz A_{iso}= 44.3 ± 0.2 MHz	33 ± 5 (7.1 ± 0.7 s) (10 mM)		
bPyTol (256 g·mol^{-1}) (**9**)		 W	A_{iso}= 46.7 ± 0.2 MHz		26 ± 1 (2.8 ± 0.1 s) (32 mM)	
O-CD$_3$bPyTol (324 g·mol^{-1}) (**10**)		 W	A_{iso}= 46.4 ± 0.2 MHz		42 ± 1 (2.4 ± 0.2 s) (32 mM)	
O-MbPyTol (312 g·mol^{-1}) (**11**)		 W	A_{iso}= 46.4 ± 0.2 MHz	119 ± 25 (9.6 ± 1.0 s) (20 mM)	38 ± 1 (1.7 ± 0.1 s) (32 mM)	73 ± 2 (12.0 ± 1.2 s) (20 mM)
C-MbPyTol (312 g·mol^{-1}) (**12**)		 W	A_{iso}= 44.7 ± 0.2 MHz	41 ± 1 (7.0 ± 1.0 s) (20 mM)	17 ± 1 (1.5 ± 0.1 s) (32 mM)	46 ± 3 (11.2 ± 1.0 s) (20 mM)
C-MPhTO (554 g·mol^{-1}) (**13**)		 ACN	A_{iso}= 41.3 ± 0.2 MHz		91 ± 1 (2.5± 0.7 s) (32 mM)	
O/C-MPhTO (554 g·mol^{-1}) (**14**)		 ACN	A_{iso}= 41.1 ± 0.2 MHz		107 ± 2 (2.3 ± 0.7 s) (32 mM)	

Figure 1: The radicals investigated in this work. Abbreviations and molecular weights (column 1), molecular structures (column 2), Room-temperature X-band EPR spectra (~0.1 mM in TCE, w (water), ACN (acetonitrile) or D (DMSO-d_6:D_2O:H_2O 60:30:10 v/v/v) (column 3), Room-temperature X-band EPR exchange coupling J, corresponding distribution $\sigma(J)$ and isotropic hyperfine A_{iso} parameters extracted from fitting analysis detailed in SI 10. An exchange coupling Hamiltonian $J\,\boldsymbol{S}_1{\cdot}\boldsymbol{S}_2$ was used. (Note that another convention exists that leads to only half the value for $|J|$.) See SI for a discussion of the sign (column 4). MAS DNP enhancements, polarization build-up times (in black) and concentration (in blue) measured at 9.4 T, 100 K in DMSO-d_6: D_2O:H_2O (60:30:10 v/v/v) (column 5), in TCE (column 6) and glycerol-d_8:D_2O:H_2O (60:30:10 v/v/v) (column 7). PEG = (CH_2CH_2O). The prefixes O- (green), C- (red) and O/C- (dark green) indicate open, closed and open/closed conformers respectively. *EPR parameters, MAS DNP enhancement and build-up time from ref.[54]. A confidence interval of about 10% for enhancement and build-up time is assumed.

Interconversion between the two conformers in AMUPol is rapid at room temperature, and the two conformers cannot be isolated. In the new bi-radicals ((**2**)-(**8**)), the conformation of the tetrahydropyran rings is locked by the introduction of *cis* CH_3 groups in positions 2 and 6. The conformation with two synaxial methyl groups is energetically unfavorable and does not contribute appreciably, *therefore only the conformations with the two methyl groups in the equatorial positions are observed.* By fixing the absolute configuration of the asymmetric carbon atoms 2- and 6-, it is possible to fix the conformation of the tetrahydropyran (THP) rings in the bi-radicals (**2**)-(**8**), generating molecules with open (O-) and closed (C-) conformations around the nitroxide moiety. In particular, we synthetized and isolated three diastereoisomeric meso forms: the first, with the 2 and 6 THP positions with configuration *S, R, S, R,* has the two THP rings in an open conformation (O-), conversely, when the configuration is *R, S, R, S* a second diasteroisomer with the THP rings in a closed conformation (C-) is generated. The third diasteroisomer synthetized has a configuration *S, R, R, S,* or (*R, S, S, R*) with the two THP rings one in the open and the other in the closed conformation (O/C-). Following synthesis, the different O- and C- conformers can be separated by silica chromatography and the conformations were determined by X-ray structure analysis (Figure S16-S19). There is no conformational exchange between the O- and C- conformers because they are diastereomers, as shown by liquid-state NMR spectra of their reduced forms (see SI for details). We have synthesized O- and C- di-nitroxides with two different types of backbone (AMUPol: (**2**)-(**6**), PyPol:(**7**), (**8**)) and different lengths of PEG chains (4 units: (**2**)-(**4**), 2 units: (**5**)-(**6**)). We further demonstrate the validity of the new design principle by investigating mono-nitroxides with either methyl ((**11**)-(**12**)) or phenyl ((**13**)-(**14**)) groups as locking substituents. The possible role of deuterated methyl groups is considered by comparing (**10**) to (**11**).

Room temperature EPR characterization

In order to characterize the local structure of the paramagnetic centers in the new radicals, we performed room temperature X-band (9.4 GHz, 0.35 T) CW EPR experiments of 100 μM solutions in either TCE, water or acetonitrile as reported in Figure 1 (third column). All of the biradicals show a complex splitting pattern in solution resulting from an interplay of hyperfine coupling to ^{14}N of the nitroxide (A_{iso}) and the exchange interaction (J) between the two unpaired electrons. The solution-state EPR data in Fig. 1, were fitted as described in Section SI 10, assuming a Gaussian distribution (with standard deviation $\sigma(J)$) of exchange couplings and the exchange Hamiltonian expressed as $J \cdot S_1 \cdot S_2$. (We note that another convention exists for which the Exchange Hamiltonian is $-2J \cdot S_1 \cdot S_2$ resulting in only half of the value for $|J|$[60] (see section SI 10).)

The EPR spectrum of O-MAMUPol (**2**) can be fitted with A_{iso}~44.7 MHz, and $|J|$ =48.0 ± 2.0 MHz, and the width of the J coupling distribution, $\sigma(J)$, of about 8 MHz. It is characterized by $|J| \approx A_{iso}$, in which case up to 15 EPR lines could be observed[54].

For C-MAMUPol (**3**), the spectrum can be fitted with A_{iso}~43.0 MHz, and with $|J|$=0 MHz, and $\sigma(J)$=150 MHz. Molecular dynamics simulations in SI 9 also show a different distribution of exchange couplings between (**2**) and (**3**), with conformations characterized by $|J| > A_{iso}$ in (**3**) that are absent for (**2**). Low temperature ED EPR spectra are also consistent with weaker and better-defined J-coupling for O-MAMUPol (**2**) and stronger and broadly distributed J-couplings for C-MAMUPol (**3**) (see section SI 10).

For OC-MAMUPol (**4**) the fit yields A_{iso}=44.7 MHz, $|J|$=110.0 ± 5 MHz and $\sigma(J)$=12 MHz. This bi-radical is characterized by $|J| > A_{iso}$.

HydrOPol (**5**) (A_{iso}=44.3 MHz, $|J|$=42.0 ± 2 MHz and $\sigma(J)$=0.5 MHz) is in a similar regime to (**2**), with $|J| \approx A_{iso}$ but characterized by a smaller distribution of J.

C-HydrOPol (**6**) (A_{iso}=44.3 MHz, $|J|$=450 ± 50 MHz and $\sigma(J) \leqq 100$ MHz) has $|J| \gg A_{iso}$ with a relatively broad distribution of J.

Similarly, O-PyPolC_6OH (**7**) (A_{iso}=46.7 MHz, $|J|$=46.6 ± 2.0 MHz, $\sigma(J)$=4 MHz) and C-PyPolC_6OH (**8**) (A_{iso}=44.3 MHz, $|J| \gtrsim 400$ MHz, $\sigma(J) \leq 100$ MHz) belong to the $|J| \approx A_{iso}$ and $|J| \gg A_{iso}$ regimes respectively.

The EPR spectra of the mono-nitroxides are all characterized by a 1:1:1 triplet resulting from the hyperfine coupling to ^{14}N with the following fitted values: bPyTol (**9**), reference for (**10**)-(**14**), A_{iso}=46.7 MHz, O-CD_3bPyTol (**10**) A_{iso}=46.4 MHz, O-MbPyTol (**11**) A_{iso}=46.4 MHz, C-MbPyTol (**12**) A_{iso}=44.7 MHz, C-MPhTO (**13**) A_{iso}=41.3 MHz and O/C-MPhTO (**14**) A_{iso}=41.1 MHz. In C-MbPyTol (**12**) each line is further split, both for the CW room temperature EPR spectrum in water (Figure 1) and in D_2O (See Table S7 and Figure S37). We attribute this additional splitting to a hyperfine interaction between the electron spin and two axial or equatorial proton spins of the central ring (section SI 10).

MAS DNP performance

We observe a stark difference in MAS DNP performance between C- and O- conformers for each pair in the series despite their otherwise identical constitution. Figure 1 shows the ^1H MAS DNP solvent enhancements ($\varepsilon_{C\,CP}$ or ε_{1H} for (**9**)-(**12**) in TCE) obtained in DMSO-d_6:D_2O:H_2O (60:30:10 v/v/v) at ~100 K, 9.4 T and 8 kHz MAS rate for the following radical concentrations: 3.2 mM (O-MAMUPol (**2**), C-MAMUPol (**3**)), 2 mM (O/C-MAMUPol (**4**)), 10 mM (HydrOPol (**5**), C-HydrOPol (**6**), O-MPyPolC_6OH (**7**) and C-MPyPolC_6OH (**8**)) and 20 mM (bPyTol (**9**), O-CD_3bPyTol (**10**), O-MbPyTol (**11**), C-MbPyTol (**12**), C-MPhTO (**13**) and O/C-MPhTO (**14**)). DMSO-d_6:D_2O:H_2O was used for comparison across the whole series because of otherwise generally lower solubility for some in glycerol-d_8:D_2O:H_2O.

The bi-radicals (**2**) and (**3**), having an identical urea-based linker and differing only in the conformation of the tetrahydropyran rings, represent the O- and C- conformers most closely related to AMUPol (**1**). MAS DNP experiments were performed at a concentration of 3.2 mM (dictated by the poor solubility of the C-conformer in the DMSO/water solvent mixture). An order of magnitude difference in the measured $\varepsilon_{C\,CP}$ is observed: 250 ± 20 for (**2**) and 23.0 ± 2.5 for (**3**) (see Figure 1 column 5). For comparison, AMUPol (**1**) at 3.2 mM in the same solvent mixture, field and temperature yields an $\varepsilon_{C\,CP}$ = 175 ± 20 (and at 10 mM the enhancement we obtain is $\varepsilon_{C\,CP}$ = 183 ± 20). In addition, O/C-MAMUPol (**4**) yields at 2 mM an $\varepsilon_{C\,CP}$=211 ± 16 (see Figure 1 column 5).

In order to address whether this difference in DNP enhancement is solvent specific, we have investigated alternative DNP matrices (Figure 1 column 6-7). All solvent mixtures investigated are routinely used in MAS DNP experiments and known to form good glass at cryogenic temperatures.[79] Specifically, (**2**) and (**3**) have been tested at 16 mM, optimal concentration for the TEKPol radical family[55], in TCE[79] returning $\varepsilon_{C\,CP}$ = 102 ± 20 for (**2**) and $\varepsilon_{C\,CP}$ = 14 ± 2 for (**3**), confirming the difference of almost one order of magnitude in the MAS DNP results between O- and C- conformers. O/C-MAMUPol (**4**) at 16 mM in TCE yields $\varepsilon_{C\,CP}$=78 ± 3 which is in between the value for (**2**) and (**3**). It has been reported that the overall ^{1}H concentration in TCE and in glycerol-d_8:D_2O:H_2O (60:30:10 v/v/v) or DMSO-d_6:D_2O:H_2O (60:30:10 v/v/v) is identical[79]. As reported in Table S2, at 21.15 T, 16 mM of (**4**) in TCE shows a better enhancement than the widely used TEKPol at the same concentration (9.7 ± 1.0 versus 6.5 ± 0.5). The ~50% difference in enhancement between DMSO-d_6:D_2O:H_2O (60:30:10 v/v/v) and TCE for both (**2**) and (**3**) can be due to the different conformations assumed by the radicals in different solvents, but also to the tendency of urea moieties to interact through complementary hydrogen bonds, thus potentially favoring aggregation processes in TCE as described for the urea-based series in ref.[55]

HydrOPol (**5**) is the open methylated conformer equivalent of PyPolPEG$_2$OH which yielded $\varepsilon_{C\,CP}$=303 at 10 mM in glycerol-d_8:D_2O:H_2O (60:30:10 v/v/v) in reference[54]. Here (**5**) dissolved in DMSO-d_6:D_2O:H_2O (60:30:10 v/v/v) at 10 mM and measured under otherwise identical field and temperature conditions results in a similar proton enhancement $\varepsilon_{C\,CP}$=330. The closed analogue C-HydrOPol (**6**) dissolved at 10 mM in identical DMSO/water solvent yielded an order of magnitude smaller enhancement ($\varepsilon_{C\,CP}$=29 ± 3) at 9.4 T and 100 K. The corresponding Zeeman field profile for (**5**) at 9.4 T and 100 K is reported in Fig. S29. The difference between positive and negative DNP peaks is approximately 0.025 T corresponding, in electron frequency units, to 1.57×ν_H, with ν_H the ^{1}H Larmor frequency at 9.4 T. We have also conducted experiments on (**5**) at 5 mM using a 1.3 mm MAS DNP probe enabling fast MAS up to 40 kHz under cryogenic DNP conditions. In this case a proton enhancement of ε_{1H}=293 at 8 kHz and 9.4 T was measured (see Table S3 and Figure S31). This enhancement was further improved to ε_{1H}=350 upon addition of manually ground sapphire particles[84].

O-MPyPolC$_6$OH (**7**) and C-MPyPolC$_6$OH (**8**) are the O- and C-conformers of PyPolC$_6$OH which yielded $\varepsilon_{C\,CP}$=290 in glycerol-d_8:D_2O:H_2O (60:30:10 v/v/v) at 10 mM in reference [54]. Here (**7**) and (**8**) dissolved in DMSO-d_6:D_2O:H_2O (60:30:10 v/v/v) at 10 mM and measured under otherwise identical field and temperature conditions yield $\varepsilon_{C\,CP}$=203 ± 15 and $\varepsilon_{C\,CP}$=33 ± 5 respectively. The results for (**7**) and (**8**) correlate with the previous observation for (**2**) and (**3**) in the same solvent (see Figure 1), confirming an almost one order of magnitude MAS DNP enhancement difference between the O- and C- classes.

A common adverse feature to many PAs is the reduced overall contribution to the NMR signal intensity due to paramagnetic bleaching caused by the PA itself.[67-68, 85] In some cases, the bleaching is aggravated as the MAS rate is increased, leading to the so-called depolarization curve.[63, 67] We have determined the contribution factors at static and 12 kHz MAS rates using a 3.2 mm sapphire rotor for some of the investigated radicals ((**1**), (**5**), (**11**), (**12**)) in DMSO/water mixture at 9.4 T and 100 K. The values are detailed in Table S4. The contribution factor for biradicals (**1**) and (**5**) in a static sample, θ_s, ranges in between 0.87 and 0.88, whereas at 12 kHz, θ_{12}, it ranges from 0.51 and 0.44. For the mono-radicals (**11**) and (**12**) at 20 mM, θ_s is 0.91 and 0.95, with θ_{12} of 0.41 and 0.52 respectively. The decreasing value of θ with the MAS rate is indicative of depolarization effects as expected for this type of nitroxide radicals if the DNP mechanism is CE.[67] In the case of mono-radicals (**11**) and (**12**) at 20 mM, the reduction of the contribution factor with the MAS rate is rationalized in terms of inter-molecular CE as discussed below.

It has recently been shown that the interplay between dipolar and exchange interactions, together with the mutual *g*-tensor orientation, is crucial for the PA efficiency, with the balance between dipolar and exchange coupling being predicted to play an important role at high magnetic fields.[53, 59-61, 73]

We have carried out MD simulations in explicit water, as detailed in SI 9, in order to estimate the relative orientations of the nitroxide moieties in open and closed forms. As expected for the linkers used here, the mean angles observed in the simulations for (**2**), (**3**) and (**7**) and (**8**) between the NO tensors excludes any collinearity, that would result in no DNP. Figure S36 shows only a small change in the relative *g* tensor orientations between O- and C- isomers (e.g. from $\langle\theta\rangle$ = 138 ± 11° for O-MAMUPol(**2**) (Fig. S36a)to $\langle\theta\rangle$ = 149 ± 12° for C-MAMUPol(**3**) (Fig. S36b) and from $\langle\theta\rangle$ = 126 ± 11° for O-PyPolC$_6$OH (**7**) (Fig. S36e) to $\langle\theta\rangle$ = 140 ± 13° for C-PyPolC$_6$OH (**8**) (Fig. S36f)), which does not correlate with the large MAS DNP differences experimentally observed.

As regards the performace at higher fields, at 21.15 T the ^{1}H enhancement of (**5**) at 10 mM drops, as expected, to around 10. Noticeably, as reported in Fig. S30 and Table S2, at 21.15 T, 8 kHz MAS and in DMSO-d_6:D_2O:H_2O (60:30:10 v/v/v), 2 mM HydrOPol (**5**) yields a factor ~4 better MAS DNP enhancement than C-HydrOPol (**6**) (21.6 ± 5.5 vs 4.3 ± 0.6). The better performance reduces to a factor of ~3 (9.7 ± 1.0 vs 3.2 ± 0.3) at 10 mM under the same conditions of solvent and magnetic field. This similar behavior would suggest that there is no big change in the magnetic interactions ~~between~~ for (**5**) and (**6**) in frozen solution.

AMUPol (**1**) in glycerol/water mixtures, for example, yields proton enhancements that change from 20 to 250 in between 21.1 T and 9.4 T. We note that M-TinyPol, the O- analogue of TinyPol, has been recently shown a remarkable proton enhancement of 32 and 90 at 10 mM, 100 K at 21.1 T and 18.8 T respectively. M-TinyPol and AMUPol (**1**) have similar exchange values, i.e. $|J|\approx$30 MHz, but different dipolar interactions[61].

In the bi-radical series ((**1**)-(**8**)) the measured exchange interaction from room temperature solution-state CW X band EPR for two of the closed radicals ((**6**) and (**8**)) is up to an order of magnitude higher than that for the open analogues. Specifically, for C-MAMUPOL (**3**) and O-MAMUPOL (**2**) the measured *J* value changes from a distribution of 0 ± 50 MHz to 48 ± 2 MHz,

respectively. In contrast, for HydrOPol (**5**) and C-HydrOPol (**6**) the measured *J* value changes from a distribution of 42 ± 2 MHz to 450 ± 50 MHz! The result is similar for the pair ((**7**), (**8**)). However, these room temperature solution-state values should be interpreted with caution. In solution at room temperature, it is possible that the conformational dynamics leads to rare instances where the two nitroxide groups are in very close proximity, which could lead to extremely large instantaneous *J* values, and large average *J* values. The sterics of the open- and closed- forms are such that this is much less likely to occur in the open isomers. This could lead to a very large average *J* coupling in solution, which would not be representative of the conformations present in frozen solutions. ~~In particular,~~ we note that also ~~the~~ MD simulations (see Figure S34) do not predict, for O-MAMUPol (**3**) and C-MAMUPol (**2**), such large conformational changes that would lead to these extreme *J* values. In Figure S34 we see a modest shift to slightly higher *J* couplings in the closed isomers, and to a slightly broader distribution of couplings.

The usually reported matching condition for CE, e.g. ω_{01}- ω_{02}=± ω_{0H} with ω_{01}, ω_{02} electron frequencies and ω_{0H} proton nuclear Larmor frequency, assumes that D=-$(d+J)$ ≪ ω_{0H} with d and J being the electron-electron dipole and exchange interactions respectively. The d value for AMUPol (**1**) is of the order of 30 MHz, and is presumably similar for the other bi-radicals here (see for example Table S6 and Fig. S36c and S36d for O-MAMUPol (**2**) and C-MAMUPol (**3**) and Fig. S36g and S36h for O-PyPolC$_6$OH (**7**) and C-PyPolC$_6$OH (**8**) extracted from the MD trajectories). The average solution-state |J| value for the C-radical (**6**) and (**8**) is estimated from room temperature EPR data to be the same order of magnitude as ω_{0H} at 9.4 T and would then be predicted to yield relatively much better performance at higher magnetic fields. However, as shown in Fig. S30 and Table S2, HydrOPol (**5**) at 21.15 T and at different concentrations still performs better than C-HydrOPol (**6**).

As a result, we conclude that the *J* couplings in frozen solution are more likely to be in line with the predictions from the MD simulations, and that the very large measured solution-state values are not relevant to the DNP conditions here. Measuring the *J* values and their distributions through a multi-field EPR study is out of the scope here, and will be the subject of further work.

The changes in the *J* couplings for (**6**) and (**8**) could nevertheless partially explain its relatively poorer MAS DNP results at 9.4 T.

Here, in order to exclude that the difference in the MAS DNP enhancement between O- and C- conformers is exclusively *J*-related, we have investigated the mono-radicals (**9**) to (**14**) where intramolecular *J* interactions are absent. The open mono-radical O-MbPyTol (**11**), at 20 mM, 9.4 T and 100 K, yields an exceptional $\varepsilon_{C\,CP}$ = 119 ± 25, which is a factor ~3 higher than that of the closed analogue C-MbPyTol (**12**), $\varepsilon_{C\,CP}$ = 41 ± 1 in DMSO/water mixture. We note that this is by far the highest enhancement so far reported for a mono-nitroxide under these conditions.

The trend for mono-radicals is confirmed as O-MbPyTol (**11**), at 16 mM in TCE (see also Figure S32 for sample temperature), yields an ε_{1H} = 38 ± 1, and yields 73 ± 2 at 20 mM in a glycerol/water mixture, whereas C-MbPyTol (**12**), at 16 mM in TCE, yields an ε_{1H} = 17 ± 1, and yields 46 ± 3 at 20 mM in a glycerol/water mixture.

Remarkably, despite the absence of exchange coupling, a significant difference in MAS DNP efficiency between O- and C- persists across different solvents for the mono-radicals. We note that the factor of ~2 to 3 between (**11**) and (**12**), in different solvents, does not match the factor of ~10 between (**2**) and (**3**), (**5**) and (**6**) or between (**7**) and (**8**). For mono-nitroxides the DNP mechanism is expected to be mostly SE at low radical concentration, and it is expected to have an increasing contribution from inter-molecular CE with increasing the radical concentration.

We have subsequently tested the concentration dependence of the DNP enhancement for (**11**) and (**12**). Figure 2 shows the results in the concentration range between 1 and 40 mM in DMSO-d_6:D$_2$O:H$_2$O (60:30:10 v/v/v). While there are no major differences in the polarization build up times between the two classes (except between 1 and 3 mM), we observe a significantly higher $\varepsilon_{C\,CP}$ for (**10**) than for (**11**) at all concentrations over 1 mM. The enhancements level off for concentrations equal or greater than 20 mM. The combination of SE and CE DNP mechanisms can account for the observed trend of the enhancement as a function of the concentration with the relative weight of intermolecular CE increasing at increased radical concentrations and possibly producing a larger difference in MAS DNP enhancements between (**11**) and (**12**). Both radicals exhibit the expected concentration dependence, analogous to that previously observed for di-nitroxides[86]. Finally, the different combination of DNP mechanisms could account for the factor of ~2-3 difference (e.g. 119/41 for (**11**) and (**12**) versus 250/23 for (**2**) and (**3**)) between mono-radicals and bi-radicals in MAS DNP performance.

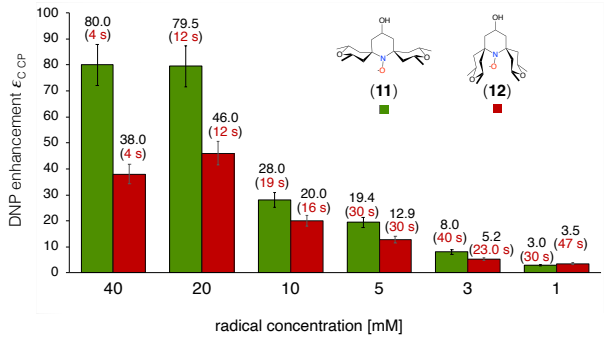


Figure 2: ^{1}H MAS-DNP enhancements measured via ^{1}H-^{13}C CP for the mono-radicals O-MbPyTol ((**11**) in green) and C-MbPyTol ((**12**) in red) in DMSO-d_6:D$_2$O:H$_2$O (60:30:10 v/v/v) at different radical concentrations at 9.4 T and 100 K. The observed $\varepsilon_{C\,CP}$ value is indicated in black above the bar. The build-up time of the solvent ^{1}H signals is given in parentheses. A 10% confidence interval on the measured enhancement is indicated.

Are C-radicals useless?

Despite a comparatively poorer enhancement for all of the investigated C- radicals ((**3**), (**6**), (**8**) and (**12**)) they may still be interesting candidates in applications performed under reducing conditions (in-cell DNP[2, 87] for example). In AMUPol (**1**) the room temperature dynamic process exchanging O- and C- conformations will lead, in presence of reducing agents, to the quick inactivation of the radical.[65] The same process will be operational for both the O- and C- conformers. However, due to the different degree of steric hindrance, the C- conformer is inactivated on a longer timescale. Figure 3 shows the reduction kinetics, measured by EPR, for the mono-radicals (**9**), (**11**) and (**12**) with 0.2 mM of ascorbic acid in H_2O at 295 K. The slower reduction for C-MbPyTol (**12**) is indicative of a better resistance to reducing environments. We ascribe this effect to different solvent accessibilities to the nitroxide regions, caused by different local conformations, as further supported below by ESEEM experiments and calculated solvent accessibility surface (Fig. S33).

In addition to the interplay between different DNP mechanisms, the clear-cut difference in terms of MAS DNP enhancement between the O- and C- classes shown in Figure 1-2 could be due to a number of further factors: (a) differences in electron relaxation times, (b) influence of local relaxation sinks, (c) other factors induced by the structural difference between the conformers. Importantly, the relative difference in performance persists across different solvents and concentrations, and is observed in both mono- and bi-radicals.

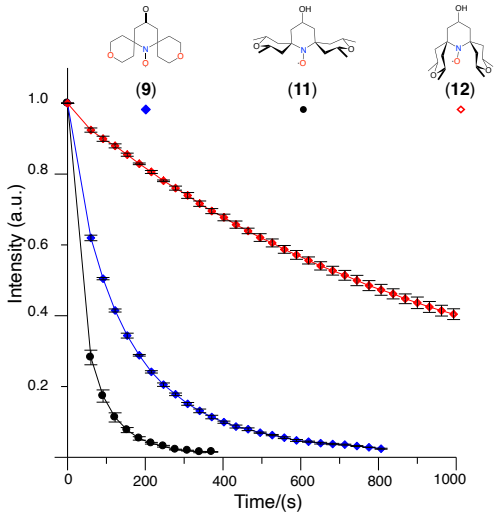


Figure 3: Reduction profiles measured through the intensity of EPR signals of 0.2 mM nitroxides ((**9**) in blue, (**11**) in black and (**12**) in red) in 0.2 mM of Ascorbic acid in H_2O at 295 K.

Saturation Factor

Previous investigations showed that MAS DNP enhancement is closely related to the electron spin relaxation properties of the radical, and the phenomenological *saturation factor* (T_{1e}·T_{2e} or T_{IR}·T_M) is used as a convenient metric, with higher saturation factors leading to larger MAS DNP enhancements[55, 74]. In the case of (**2**) and (**3**) pulsed EPR experiments at 94.1 GHz were conducted at 3.2 mM in DMSO-d_6:D_2O:H_2O (60:30:10 v/v/v) and 16 mM in TCE and 105 K in order to determine electron T_{IR} and T_M (Section SI10). The results are summarized in Table 1. In DMSO/water there is virtually no difference between the saturation factors of (**2**) and (**3**). In TCE, the saturation factor of C-MAMUPol (**3**) is a factor two larger, which does not correlate with its much lower MAS DNP enhancement. This surprising result indicates that *the electronic relaxation properties do not explain the difference in MAS DNP performance between O- and C- conformers.*

Table 1: Saturation factors ($T_{IR}T_M$ [(μs)²]) and MAS DNP enhancements ($\varepsilon_{ON/OFF}$) for a 3.2 mM solution of **2** and **3** in D/w (DMSO-d_6:D_2O:H_2O 60:30:10 v/v/v) and for a 16 mM solution of **2** and **3** in TCE (1,1,2,2-tetrachloroethane) at 105 K.

	T_{IR} [μs]		T_M [ns]		$T_{IR}T_M$ [(μs)²]		$\varepsilon_{ON/OFF}$	
R	D/w	TCE	D/w	TCE	D/w	TCE	D/w	TCE
2	168	78	581	411	98	32	250	102
3	116	70	862	864	100	81	23	14

The role of -CH_3 groups

The electronic phase memory time, T_M, in glassy frozen solutions at 100 K, is largely driven by molecular motions and librations,[88] and rotation of methyl groups is effective at inducing transverse electron relaxation. The efficiency of this process is a function of temperature and depends on the energy barrier for methyl rotation.[88] The presence of methyl groups in the proximity of the unpaired electron therefore acts as an electronic relaxation sink and is thought to hamper MAS DNP performance. Methyl groups, which still undergo fast rotation at 100 K, also act as nuclear relaxation sinks, and this is known to be detrimental for MAS DNP[79]. Indeed, deuteration of the methyl locking groups in (**10**) results in a ~10 % greater enhancement as compared to (**11**) (ε_{1H} = 42 ± 1 for (**10**) versus ε_{1H} = 38 ± 1 for (**11**)) and ~40% longer polarization build-up time ($T_{b,on}$=2.4s for (**10**) versus $T_{b,on}$=1.7s for (**11**)) in TCE. It is also worth mentioning that for nuclear spins DNP-induced cross-relaxation in methyl groups is an interesting feature for protein and amino acids investigations[89-90] and have been shown to have long-lived properties, in specific cases, at about 1 K[91-92]. Table 1 shows the electron relaxation data obtained for (**2**) and (**3**) under conditions identical to those used in a MAS DNP experiment. The T_M of C-MAMUPol (**3**) is longer than that of O-MAMUPol (**2**) in both DMSO-water and TCE, even though in the C- conformer the methyl groups are closer in space to the unpaired electron. The distribution of distances between the unpaired electron and the methyl group protons is shown in Figure 4b, extracted from the MD trajectories in section SI 9, and we find values between 4 and 5.5 Å for (**12**), whereas for (**11**) we find ~4.5 Å and ~6.3 Å respectively (see also Fig S18 and Table S10).

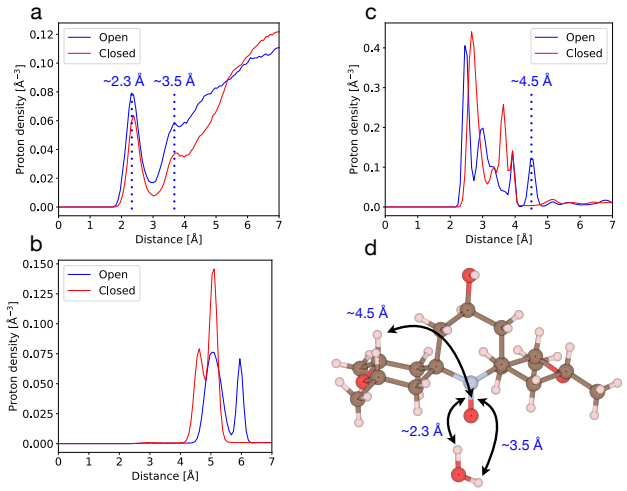


Figure 4: Proton densities for solvent water protons (a), methyl CH_3 (b) and other radical protons (c) in a sphere of 7 Å from the NO radical from MD trajectories described in section SI 9. In (d) pictorial representation of the involved nuclei and distances to NO considering for simplicity the structure of O-MbPyTol (**11**)

This result shows that differences in the methyl-induced electron relaxation do not appear to be the limiting factor in determining the differences in DNP performance between the open and closed radicals. Additionally, a comparison between the MAS DNP enhancements of (**13**) and (**14**), where methyl groups have been replaced by bulkier phenyl substituents, again shows better performance for the O/C- conformer, despite a similar build-up time for proton polarization in TCE.

Correlation between solvent accessibility and MAS-DNP enhancement

We have hypothesized that the difference between the O- and C-conformations may affect solvent accessibility around the nitroxide, thereby affecting the first few polarization transfer steps.

The structures of the O- and C- isomers of the nitroxide were confirmed by X-ray crystal structure analysis (Fig. S16-S19). The accessibility around the N-O bond has been estimated by calculating the Solvent Accessible Surface (SAS) for a water molecule using Jmol software[94](Fig. S33). The major electron spin density of a TEMPO-like nitroxide locates on N–O site with a nearly equal distribution on the nitrogen and oxygen atoms. Values of 20.60 and 0.45 $Å^2$ for the Oxygen and Nitrogen atoms respectively for O-MbPyTol (**11**) and of 9.69 and 0.06 $Å^2$ for the Oxygen and Nitrogen atoms respectively for C-MbPyTol (**12**) were obtained, illustrating the significant difference in solvent accessibility.

Experimentally, we have probed solvent accessibility using pulsed EPR. Electron Spin Echo Envelope Modulation (ESEEM) has been demonstrated to provide reliable information on water accessibility in large transmembrane proteins[82] and water penetration in micelles by quantification of hyperfine couplings to deuterium nuclei in deuterated water molecules. It allows the estimation of the number of nuclear spins in the vicinity of the electron spin on a length-scale between 3-6 Å. The ESEEM effect arises from partially allowed transitions occurring with the simultaneous change of the electron and nuclear magnetic spin quantum numbers.

The three-pulse ESEEM consists of the following sequence of pulses: $\pi/2$-τ- $\pi/2$-T- $\pi/2$-τ-echo as described in the *EPR Spectroscopy* section above. The intensity of the echo signal is modulated by the hyperfine interaction between the electron spin and vicinal ^{2}H nuclei during the variable delay T resulting in an oscillating decay as shown in Figure 5a-c. In the three-pulse ESEEM experiment, the modulation depth, K, is defined as the peak-to-peak distance between the first maximum and the first minimum in the deuterium modulation. The ^{2}H modulation profiles for the investigated mono- and bi-radicals are shown in Figure 5a-c.

Experimentally, in order to minimize the influence of ^{1}H on the ^{2}H modulations, the measurements were performed with τ=344 ns which corresponds to the j=5 blind spot of proton modulation: $\nu_H\tau$=5, where ν_H represents the ^{1}H Larmor frequency at X band. It was previously demonstrated that the choice of j=5 leads to optimal suppression of the hydrogen-bonded deuterons and as a result to more stable fits of the non-hydrogen-bonded ESEEM modulation[82]. Fourier transformation of the normalized nuclear modulation function and computation of the absolute value provides the magnitude spectrum (see Figure 5). The intensity $I(\nu_D)$ of the ^{2}H at the frequency ν_D is proportional to the modulation depth K as discussed in [82]. The ESEEM experiments return the highest $I(\nu_D)$ of the ^{2}H signal for the O- class of bi-radical systems (Figure 5).

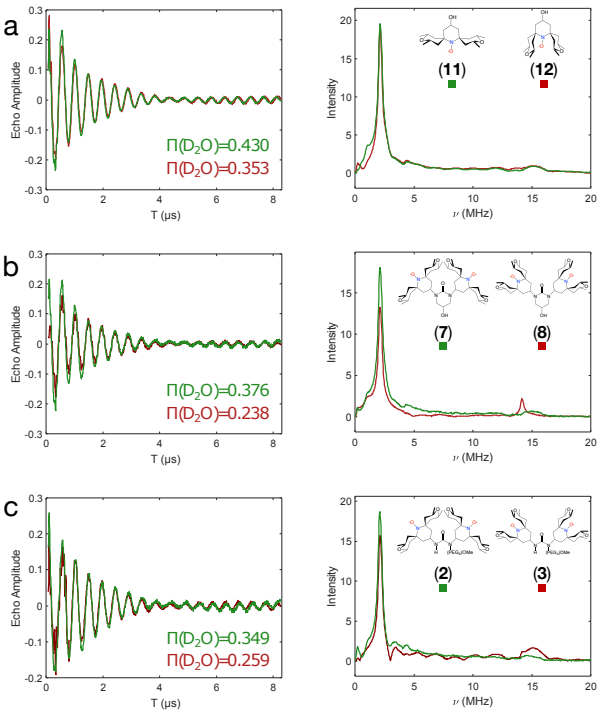


Figure 5: Three-pulse ESEEM time-domain data and the corresponding magnitude spectra in (a) for (**11**) and (**12**) in DMSO-d_6:D_2O:H_2O (60:30:10 v/v/v); in (b) for (**7**) and (**8**) in DMSO-d_6:D_2O:H_2O (60:30:10 v/v/v); in (c) for (**2**) and (**3**) in glycerol-d_8:D_2O:H_2O (60:30:10 v/v/v) at 50 K and 200 μM. The solvent accessibility parameter $\Pi(D_2O)$, defined in eq. SI-6, is always larger for the O- class: $\Pi^{O\text{-}}(D_2O) > \Pi^{C\text{-}}(D_2O)$. We note that in the case of C-

MPyPolC$_6$OH (**8**) the proton modulation is not completely suppressed leading to a small peak at the ^{1}H Larmor frequency at circa 15 MHz.

The higher intensity I(ν_D) for the O- class can be interpreted as a greater solvent accessibility to the nitroxide region for this class of PAs. The solvent accessibility parameter $\prod$, defined in [82], is 18% greater for O-MbPyTol (**11**), 37% for O-MPyPolC$_6$OH (**7**) and 26% for O-MAMUPol (**2**) as compared to the corresponding C-class analogues ((**12**), (**8**) and (**3**) respectively) and positively correlates with the MAS DNP enhancements reported in Figure 1.

In further support of modified solvent accessibility, Figure 4 shows the radial ^{1}H density in a sphere of 7 Å from middle of the NO bond calculated from 1001 snapshots from the MD trajectories for (**2**) and (**3**) (see section SI 9.2). The ^{1}H densities have been divided into three groups: water protons (Figure 4a); methyl CH$_3$ protons (Figure 4b); and other radical protons (Figure 4c). Figure S35 shows these distributions superimposed in the same vertical scale, as well as for (**7**) and (**8**). Note that in the DNP samples, the H$_2$O is 75% deuterated, so the water proton density should be divided by 4. Note also that the MD trajectories were calculated in pure H$_2$O, which does not therefore take into account the potential perturbations due to the DMSO present in the DNP formulation.

As shown in Figure 4, the water protons have a higher density for O-MAMUPol (**2**) as compared to the closed analogue, which is in good agreement with the ESEEM measurements discussed above. The two local maxima at ~2.3 Å and ~3.5 Å are due to the first directly coordinated water molecule (see Figure 4d). Table S5 gives the number of the different types of protons contained in 5, 6 and 7 Å radii around the nitroxide. We note also that the CH$_3$ protons are not only closer to the NO in the closed forms, as discussed previously above, but also have a region of higher density. In addition, in O-MAMUPol (**2**) there is significant radical proton density reaching out to longer distances (around ~4.5 Å) than in the closed isomer. All these differences can lead to different polarization transfer pathways between the O- and C-radicals into the bulk. For example the lower density of CH$_3$ protons as compared to the solvent and radical protons could potentially modify the polarization dynamics in the O-isomers across the spin diffusion barrier normally present in DNP experiments[93].

CONCLUSIONS

We have introduced a set of 13 new nitroxide mono- and bi-radicals as PAs for MAS DNP at 9.4 T and 100 K. They have been grouped into O- and C- classes depending on the conformation of the tetrahydropyran rings in the regions around the NO$^\bullet$ moieties (Scheme 2). We observe a significant difference in the MAS DNP performance for the two classes, with the O- class yielding an order of magnitude better result than the corresponding C- analogues for bis-nitroxides. The strikingly higher enhancement for the O- class is preserved across different DNP media. We have shown that these results cannot be exclusively rationalized in terms of solvent-induced variations in local conformation, different saturation factors or methyl-induced relaxation. The limited DNP enhancement for the C-bi-radicals can, in some cases, be partially due to stronger magnetic interactions (large values of $|J|$). However, it cannot account for the difference between C- and O-mono-radicals, where no exchange coupling is present. SAS simulations and the ESEEM experiment confirm a higher local solvent accessibility in the O-class radicals. This positively correlates with the MAS DNP results. Changes in the local concentration of the different types of protons in the vicinity of the unpaired electron could affect the first steps of the spin diffusion process, leading to the observed changes in the MAS DNP bulk solvent enhancement.

This study highlights the, so far overlooked, crucial importance of *local* conformational changes in the determination of both magnetic and bulk solvent accessibility properties, introducing a new design principle for the synthesis of efficient radicals tailored for high field MAS DNP. HydrOPol (**5**) gives a proton enhancement at 9.4 T of 330, which is about 80% higher than AMUPol (**1**) under these conditions of field, concentration and solvent (DMSO-d_6:D$_2$O:H$_2$O (60:30:10 v/v/v)). In addition, the mono-radical O-MbPyTol (**11**) yields a proton enhancement of 119 at 9.4 T in DMSO/water which is, to the best of our knowledge, the highest MAS DNP proton enhancement at 9.4 T and 100 K so far reported for a mono-nitroxide. The relevance of the principle introduced here has already very recently been supported by the introduction of M-TinyPol radical (the O- derivative of TinyPol), designed for high-fields, and which also yielded the highest reported proton enhancements so far for MAS DNP at 100 K at 18.8 T and 21.15 T (90 and 32 at 10 mM in glycerol-d_8:D$_2$O:H$_2$O (60:30:10 v/v/v) respectively). [61]

ASSOCIATED CONTENT

Supporting Information

Supporting information includes synthetic details, experimental protocols and supporting data. The raw data reported here are available at http://www.zenodo.org/xxxxxxxx (*will be completed at time of publication*).

The Supporting Information is available free of charge on the ACS Publications website.

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Author Contributions

The manuscript was written through contributions of all authors. All authors have given approval to the final version of the manuscript.

Notes

The authors declare no conflicts of interest.

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